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| -Naming Hydrates-    Example  1) Write the name of Na2SO4 \* 10H2O.    We begin by naming the ionic compound Na2SO4 using standard ionic compound naming conventions.. Using our naming conventions we know the name of the ionic compound is sodium sulfate. Now using the table at the bottom of the page we know that the prefix we need for ten water atoms is *deca-*, now just add the prefix to hydrate to get decahydrate. So the name of this hydrate is sodium sulfate decahydrate.    2) Write the formula for calcium sulphate hexahydrate.    We start this problem by using our naming conventions for ionic compounds to find that the equation of calcium sulphate is CaSO4. Now using the chart below we see that the prefix hexa- means there are six water molecules attatched to the compound CaSO4. So our equation for this hydrate is CaSO4 \* 6H2O. |

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| Practice Problems  Write the formula for the following hydrates  1)   zinc sulphate heptahydrate  2)   copper (I) sulphite monohydrate  3)   cobalt (II) fluoride tetrahydrate  4)   lithium nitrate trihydrate  5)   sodium sulphate decahydrate  Write the compound name for the following formulas  6)   Na2SO4 \* 10H2O  7)   LiNO3 \* 3H20  8)   Cu2SO 3 \* 3H20  9)   Ca(N03) 2 \* 2H20  10)   NaCl04 \* H20  [Click Here for answers.](http://www.fargo.k12.nd.us/education/components/scrapbook/default.php?sectiondetailid=8165&sc_timestamp=1116608743) |

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| Prefixes   |  |  | | --- | --- | | Prefix | Number of Atoms | | mono- | 1 | | di- | 2 | | tri- | 3 | | tetra- | 4 | | penta- | 5 | | hexa- | 6 | | hepta- | 7 | | octa- | 8 | | nona- | 9 | | deca- | 10 | |