**Science Humor**
Q: What do you call a bond that is fond of sarcasm?

A: An ironic bond.

#### Identifying Bonds

In most cases, you can use a [table of electronegativity values](http://library.thinkquest.org/3310/lographics/electro.html) to determine the type of bonds that will occur between atoms. A simple of rule of thumb goes as follows:

1. If the electronegativity difference is between 0 and 0.2, it is probably covalent
2. If the electronegativity difference is between .2 and 1.7, it is probably polar
3. If the electronegativity difference is greater than 1.7, it is probably ionic

## Electronegativity Table

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| --- |
| H2.1 |
| Li1.0 | Be1.5 |  |  |  |  |  |  |  |  |  |  | B2.0 | C2.5 | N3.0 | O3.5 | F4.0 |
| Na0.9 | Mg1.2 |  |  |  |  |  |  |  |  |  |  | Al1.5 | Si1.8 | P2.1 | S2.5 | Cl3.0 |
| K0.8 | Ca1.0 | Sc1.3 | Ti1.5 | V1.6 | Cr1.6 | Mn1.5 | Fe1.8 | Co1.9 | Ni1.9 | Cu1.9 | Zn1.6 | Ga1.6 | Ge1.8 | As2.0 | Se2.4 | Br2.8 |
| Rb0.8 | Sr1.0 | Y1.2 | Zr1.4 | Nb1.6 | Mo1.8 | Tc1.9 | Ru2.2 | Rh2.2 | Pd2.2 | Ag1.9 | Cd1.7 | In1.7 | Sn1.8 | Sb1.9 | Te2.1 | I2.5 |
| Cs0.7 | Ba0.9 | La-Lu1.0-1.2 | Hf1.3 | Ta1.5 | W1.7 | Re1.9 | Os2.2 | Ir2.2 | Pt2.2 | Au2.4 | Hg1.9 | Tl1.8 | Pb1.9 | Bi1.9 | Po2.0 | At2.2 |
| Fr0.7 | Ra0.9 | Ac1.1 | Th1.3 | Pa1.4 | U1.4 | Np-No1.4-1.3 |  |  |  |  |  |  |  |  |  |  |